

Synthesis and Characterization of Nickel and Zinc Sulfide Ethylenediamine Complexes

S. Abd El Mongy¹, R. Seoudi² and A. A. Shabaka²

1) Physics Department, Faculty of Girls of Science and Arts, Ain Shams University, Cairo, Egypt.

2) Spectroscopy Department, Physics Division, National Research Center, Dokki, Cairo, Egypt.

In the present work, nickel and zinc sulfides ethylenediamine ($N_2H_4C_2H_4$) complexes were prepared in an autoclave at 180 °C for 12 h. Transmission electron microscope (TEM) and X-ray powder diffraction were used to study the morphology of the two samples. The infrared absorption spectra were obtained and discussed in the range (200-4000 cm^{-1}). New bands appeared at about 238,240 and 297 cm^{-1} assigned to Ni-S and Zn-S vibration coupled with metal nitrogen vibration. The data indicate that, the nickel and zinc sulfide were synthesized as complexes with ethylenediamine. Thermogravimetric analysis (TGA) shows three main regions. The first region is attributed to dehydration, the second is due to degradation and the third is due to carbonization. The electrical properties were measured in the temperature range (300-400 K), the activation energy was calculated and discussed.

1. Introduction:

Transition metal catalysts have received special attention and in particular, some metal oxides, such as TiO_2 and ZnO have been extensively studied and used as materials for environmental protection [1, 2]. Transition metal sulfide, such as CdS and ZnS , well know catalytic semiconductor. Metal complexes are known to have good catalytic properties in organic synthesis, and some systems have been used as sensitizer for photolysis of water [3, 4], but there are few examples that were used for wastewater treatment. Nickel sulfides have important applications in hydrogenation catalysis, especially in the organic synthesis procedures [5], among which $\alpha-NiS$ and Ni_3S_2 have been reported as important of catalytic compounds [6, 7]. Moreover, NiS possesses interesting electronic properties due to its metal-insulator, paramagnetic-antiferromagnetic phase change [8-10].

The present study is devoted to the preparation of ZnS and NiS ethylenediamine complexes. The obtained complexes were characterized by transmission electron microscopic, x-ray diffraction, and infrared spectroscopy. The thermal and electrical properties were measured.

2. Experimental

The solvothermal elemental reaction was used for the preparation of ZnS and NiS ethylenediamine complexes. An appropriate amount of $ZnCl_2$, $NiCl_2$ and sulfur powder (99.95%) was added to 80 ml of ethylenediamine in a Teflon-lined autoclave of 100 ml capacity. Then the autoclave was sealed and maintained at 180°C for 12 h without shaking. The precipitate was filtered off, washed with distilled water and absolute ethanol, respectively, and then dried in a vacuum at 80°C for 4h.

The microstructures and the particle distribution were investigated by Transmission electron microscope (Zeiss EM 10) operating at 100 kV. The complexes powder was prepared by making a suspension from the powder in distilled water. The suspension was sinterfuge to collimate the large size particles. Then a drop of the suspension was put into the carbon grid and left to dry at 100°C.

X-ray diffraction pattern was measured using (A Philips PW 1370) X-ray diffractometer operating at 35 kV and 15 mA, using a monochromated Ni-filter, CuK α radiation ($\lambda = 1.54$ nm), a scanning rate of 2 θ /min, and a range of $4 \leq 2\theta \leq 80$. A Jasco Model 300E Fourier Transform Infrared Spectrometer was used to measure the vibrational spectra of the investigated samples. Thermogravimetric analysis was carried out in nitrogen atmosphere, at a heating rate of 10 °C min⁻¹ using the computerized Perkin Elemer (U. S., Norwalk, C. T) TGA7-Unix system series.

The DC electrical conductivity of the samples was obtained out on samples in the form of disc of about 0.85 cm diameter and 0.16 cm thick. The two surfaces of each sample were coated with silver paint (BDH) and checked for good contact. The DC-conductivity was carried out from room temperature up to 400 K and measured using Keithly electrometer type 6517 A. The temperatures were measured by K-type thermocouple connected to Digi-Sense thermometer. The junction of thermocouple was in contact with the sample. The accuracy of measuring temperature was better than $\pm 1^\circ\text{C}$.

3. Results and Discussion:

Figure (1a, b) shows the images of transmission electron microscopy (TEM) of NiS and ZnS ethylenediamine complexes. From this figure, it is noticed that, the morphology is different for the two samples and this is due to the different formed aggregation of NiS and ZnS in the ethylenediamine. The physical and chemical properties of the two metals play an important role of this aggregation.

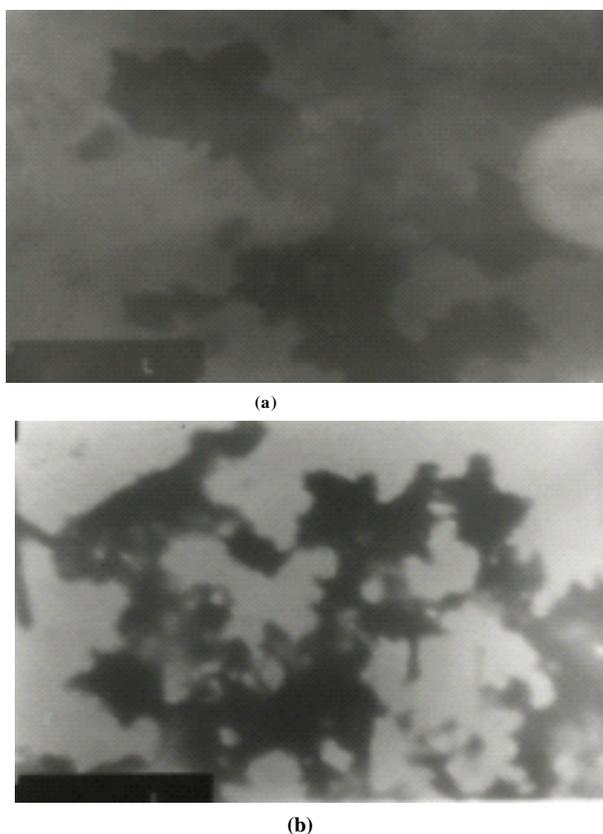
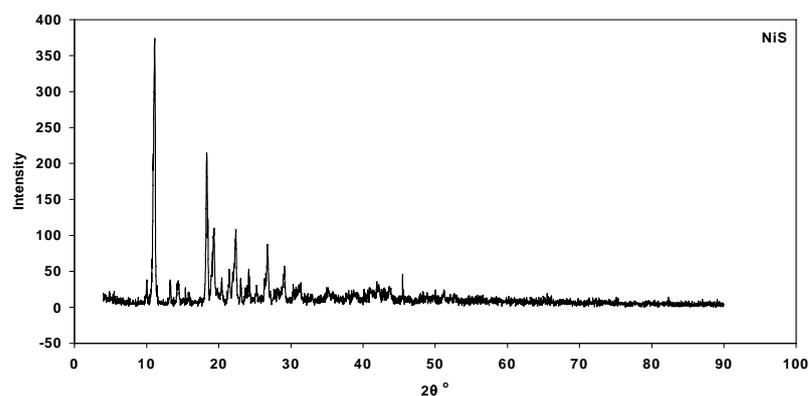
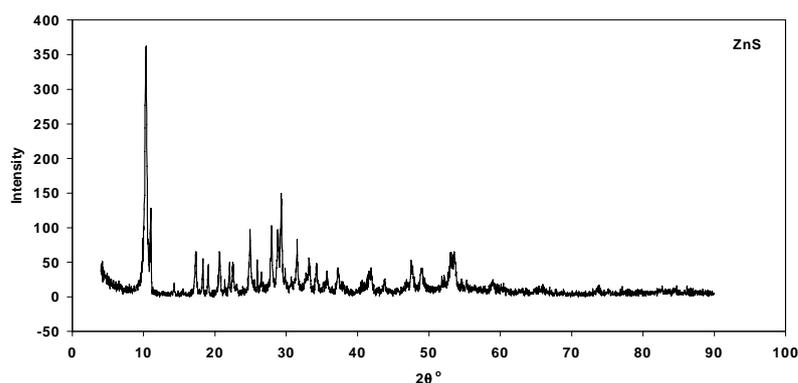


Fig. (1: a, b): Transmission electron microscope (TEM) of
(a) NiS ethylenediamine complex. (b) ZnS ethylenediamine complex.

Figure (2a, b) represent the X-ray diffraction patterns of the two samples under investigation. By comparing the d values of NiS form [JCPDS, card No., 83-0575] with the d values of NiS formed it was found that, the two values are different. Also, the d values of ZnS form [JCPDS, card No. 36-1450] disappeared in the ZnS formed. These indicate that, the two new compounds are not NiS and ZnS but it is a complexes with ethylenediamine and can be written in the form [NiS(en) and ZnS(en)].



(a)



(b)

Fig. (2): XRD pattern of (a) NiS ethylenediamine complex and (b) ZnS ethylenediamine complex.

The infrared absorption spectrum in the wavenumber range ($4000\text{-}400\text{ cm}^{-1}$) of ethylenediamine (en) is shown in Fig. (3a). From this figure it was shown that, the NH stretching vibration was appeared at 3357 and 3281 cm^{-1} . The CH symmetric and asymmetric stretching vibrations of CH_2 groups were observed at 2927 and 2855 cm^{-1} respectively. The absorption band appeared at 1598 cm^{-1} is most likely due to NH deformation vibration. To give clear information about the molecular structure changes the infrared spectrum of NiS(en) and ZnS(en) of (en) was compared with that of en. It can be noticed that, the spectra of NiS(en) and ZnS(en) are quite different than the spectrum of (en). IR absorption spectra of the prepared samples are shown in Fig. (3b, c). The NH vibration at 3357 and 3281 cm^{-1} for the (en) shifted to lower wavenumber at (3295 and 3169 cm^{-1}) in the spectrum of NiS(en) and at (3239 and 3115 cm^{-1}) in the spectrum of ZnS(en). This shift is due to the coordination

of the metal with ethylenediamine through the lone pair of electrons on the nitrogen atom. The bands appeared at 238 cm^{-1} in the spectrum of NiS Fig. (4a) due to the Ni with coordinated with the nitrogen atom. The two bands appeared at about 240 cm^{-1} and 297 cm^{-1} for ZnS Fig. (4b) may be due to coordinated Zn with nitrogen and sulfur atom.

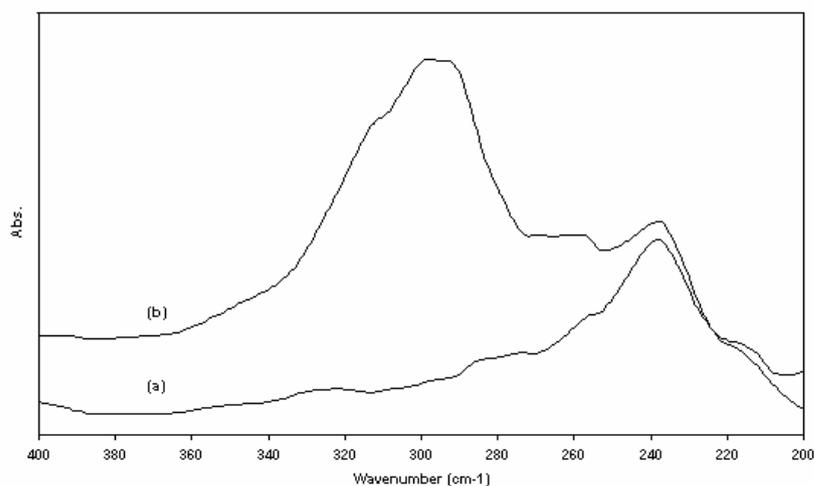


Fig.(3): Infrared absorption spectra of (a) NiS ethylenediamine complex and (b) ZnS ethylenediamine complex in the range ($400\text{-}200\text{ cm}^{-1}$)

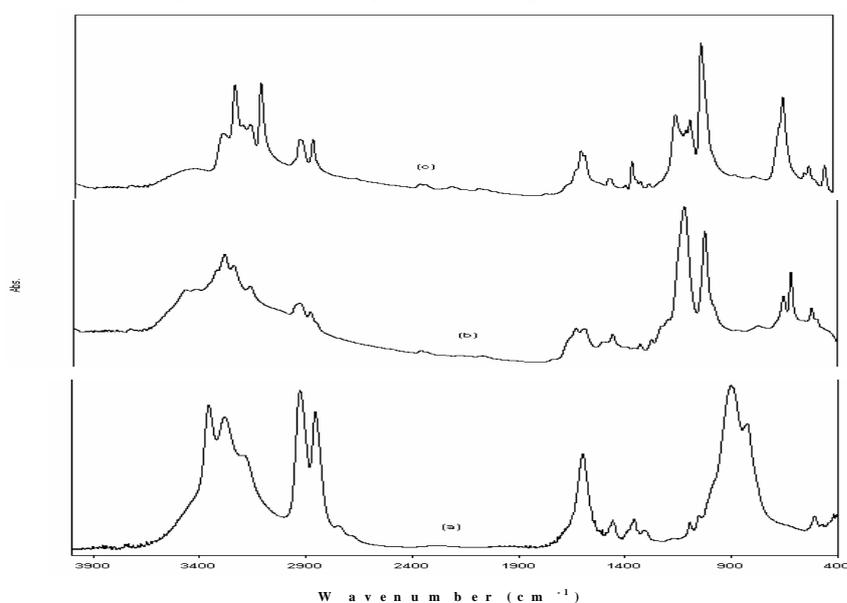


Fig.(4): Infrared absorption spectra of (a) ethylenediamine, (b) NiS ethylenediamine complex and (c) ZnS ethylenediamine complex in the range ($4000\text{-}400\text{ cm}^{-1}$)

Thermogravimetric analysis (TGA) was used to study thermal stability in the temperature range from 30 °C to 800 °C for the samples under investigation (Fig. 5). Figure (5, b) shows that more than one region for each sample are observed. The loss of weight nearly constant with temperature in the first region started from 30 °C for two samples and ended at about 100 °C for NiS(en) and ZnS(en) complexes. The second region ended at 250 and 340 °C with loss of weight 10 and 14 % for the two samples respectively. This is due to the moisture evaporation and generation of noncombustible gases such as CO, CO₂. The main decomposition region ended at about 540 °C with loss of weight 50 and 70% for the two samples respectively. This is most probably due to generation of combustible gases [11]. The last decomposition region represents the thermal degradation of the anhydride. The thermal activation energy was calculated by analyzing TGA data by differential method [12]. Plotting $\text{Log} [(w_0 - w_\infty)/(w_t - w_\infty)]$ against time (t), where w_0 is the initial weight, w_t is the weight at time t and w_∞ is the weight of the ash remaining after the final heating. As shown from Fig. (6a, b), a two linear part are noticed, the first part occurred due to the water evaporation and the second line is due to thermal decomposition. The activation energy of the main decomposition temperature is calculated using Arrhenius plot for the first order reaction [13-15]. The values of the activation energy of NiS(en) and ZnS(en) complexes were 146.6 and 225.2 KJmole⁻¹. From these two values, it is seen that the activation energy of ZnS(en) is higher than that of NiS(en). This is due to the occupied of 3d orbital of Zn metal.

The dependence of DC electrical conductivity on the reciprocal absolute temperature (1/T) may be defined using the relation $\sigma = \sigma_0 \exp^{-\Delta E/kT}$ [16], where, σ is the conductivity, σ_0 represents the value of σ at (1/T)=0, T is the absolute temperature and k Boltzman's constant and ΔE is the electrical activation energy. Two regions are observed for nickel and zinc sulfide complexes. In the first region from room temperature up to 370 k, all impurity atoms are ionized and the extrinsic carriers are not yet excited to a marked degree. This is because the density of carriers remains approximately constant and equal to the impurity concentration. Therefore, the temperature dependence on the conductivity is decided by carrier mobility. The conductivity increases with increasing temperature in the second region which started form 370 k and ended at 400 k. In this region the activation energy (0.95eV) is associated with a short-lived charge transfer between impurity and the complex. This region corresponds to an intrinsic region [11]. For zinc sulfide the conductivity nearly decrease with increasing temperature in the region (300-380 k). In this region the main scattering mechanism was scattered by thermal vibrations and the mobility decreases with rising temperature so the conductivity will decreases within this region. In the second region (380-400) the conductivity was increased with increasing temperature with activation energy (2.65 eV) and it is the same behavior of nickel sulfide.

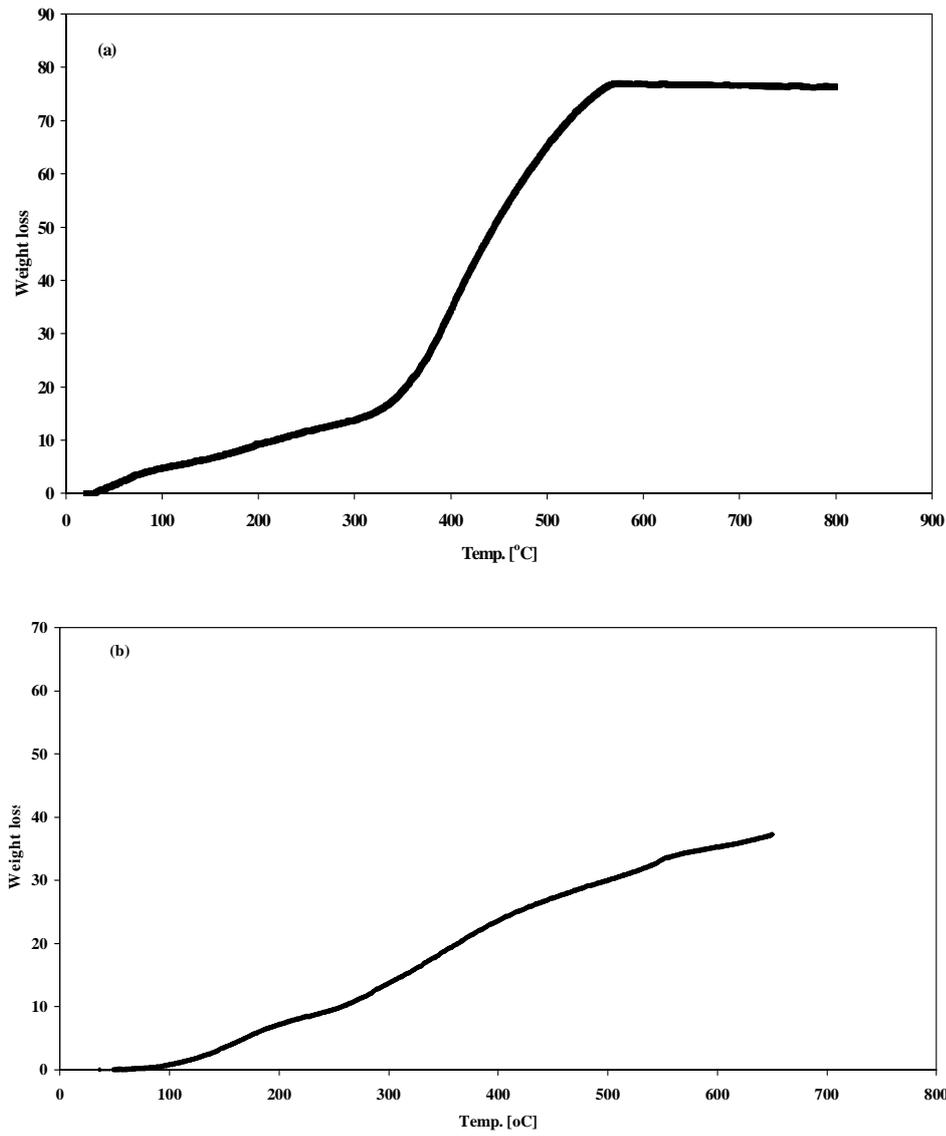


Fig.(5): Thermal gravimetric analysis of (a) NiS ethylenediamine Complex and (b) ZnS ethylenediamine complex

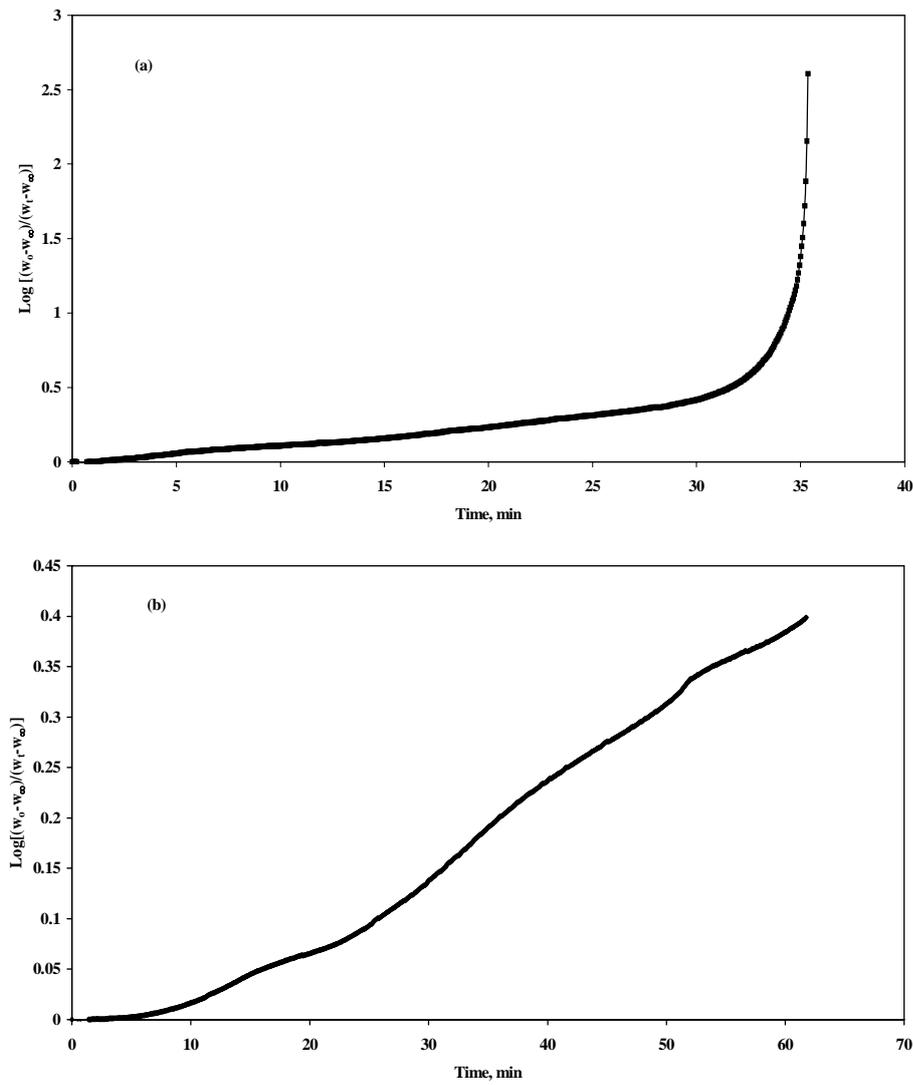


Fig. (6): Relation of $\text{Log} [(W_o - W_\infty) / (W_t - W_\infty)]$ versus time for (a) NiS ethylenediamine complex and (b) ZnS ethylenediamine complex.

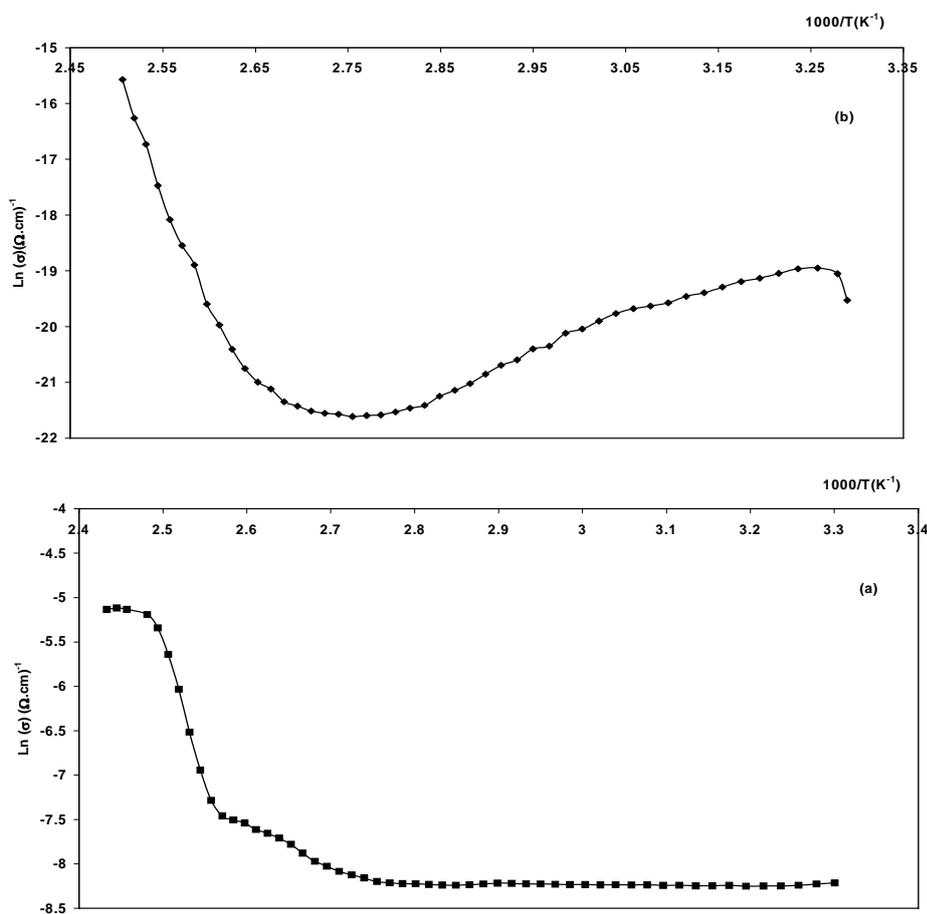


Fig. (7): Variation of DC electrical conductivity (σ) with reciprocal of absolute $1000/T(\text{K})$ for (a) NiS ethylenediamine complex and (b) ZnS ethylenediamine complex

Conclusion:

The phase and size of the nickel and zinc sulfide were studied from X-ray powder diffraction and transmission electron microscope. The results indicated that the two new compound (NiS and ZnS) complexes with ethylenediamine[NiS(en) and ZnS(en)] were formed. The complexes were confirmed by appeared of new band at about 300 cm^{-1} which assigned to Ni-S and Zn-S vibration coupled with metal nitrogen vibration. The thermal properties indicated that, the decomposition region depend on the physical properties of both Ni and Zn metal. The semiconducting behaviors were observed for two samples from DC measurements. The electrical activation energy was estimated.

References:

1. D. F. Ollis, in: H. Al-Ekabi (Ed.), "*Photocatalytic Purification and Treatment of Water and Air*", Elsevier, Amsterdam, (1993).
2. M. R. Hoffmann, S.T. Martin, W. Choi, D. W. Bahnemann, *Chem. Rev.* **95**, 69 (1995).
3. A. J. Bard, M. A. Fox, *Acc. Chem. Res.* **28**, 141 (1995).
4. M. Hara, C.C. Waraksa, J. T. Lean, B. A. Lewis, T. E. Mallouk, *J. Phy. Chem. A* **104**, 5275 (2000).
5. O. Weisser, S. Landa, "*Sulfide Catalysts, Their Properties and Applications*", Pergamon, Oxford, New York, 167 (1973).
6. Y. Yermakov, A. N. Startsev, V. A. Burmistrov, *Appl. Catal.* **11**, 1 (1984).
7. K. I. Tanaka, T. Okuhara, *Catal. Rev. –Sci. Eng.* **15**, 249 (1977).
8. J. T. Sparks, T. Komoto, *Phys. Lett.* **25A**, 398 (1967).
9. J. Trahan, R. G. Goodrich, S. F. Watkins, *Phys. Rev.* **B2**, 2859 (1970).
10. D. B. Mcwhan, M. Marezio, J.P. Remeika, P. D. Darnier, *Phys. Rev. B* **5**, 2552 (1972).
11. S. Ambily, C.S. Menon, *Thin Solid Films* **347**, 284 (1999).
12. S. Glsastone, "*Textbook of Physical Chemistry*", London, Macmilan (1962).
13. M. El-Sakhawy, *J. Therm. Anal. Cal.*, **63**, 549 (2001).
14. A. M. Kaloustain and J. Pastor, *J. Therm. Anal. Cal.*, **63** (1), 7 (2001).
15. Arthur Von Hippel, "*Dielectric Materials and Applications*", Artech House, Boston, London (1995).
16. S. I. Shihub, R. D. Gould, *Thin Solid Films*, **254**, 187 (1995).